

Please read all the questions VERY carefully before answering. On scantron start from the same bubble number as the question number for your multiple choice question. Write neatly. If I cannot read your answer, you will not receive any point. Use the attached periodic table and constant chart. No outside paper is allowed. Total points = $46 + (22 \times 3) = 112$

SHORT ANSWER. In all calculations, write the set up equation first, then put the raw data with units. Then do your calculations. Points will be deducted if your answer is not clearly written.

1) Show calculations with units and sig. fig. to convert 16.32 gallon (gal) into milliliter (mL) (given 1 gal = 3.785 L and 1 L = 1000 mL). (6 pts.) 1) _____

2) Calculate (with units and sig fig) how many in^3 are in 2.20 cm^3 (1 in = 2.54 cm.)? (8 pts.) 2) _____

3) Calculate the density of 96.0 mL of a liquid (with correct number of sig fig and units) that has a mass of 90.5 g? (6 pts.) 3) _____

4) Write the formula for (3 pts. each; Total 9 pts.):

4) _____

(a) Ammonium phosphate:

(b) Calcium Sulfate:

(c) Phosphorus pentachloride:

5) Suppose it took 108 joules of energy to raise a bar of gold from 25.0°C to 29.7°C. Given that the specific heat capacity of gold is 0.128 J/g·°C, what is the mass (in grams) of the bar of gold? Show all your calculations with set up equation and units and sig. fig.. Given $q = m.C. \Delta T$. (8 pts.)

5) _____

6) Write the name for (3 pts. each; Total 9 pts.):

6) _____

(a) $\text{Ca}(\text{HSO}_4)_2$:

(b) $\text{Al}_2(\text{CrO}_4)_3$:

(c) $\text{Co}(\text{ClO}_4)_2$:

MULTIPLE CHOICE. On scantron start from the same bubble number as the multiple choice question number. Choose the one alternative that best completes the statement or answers the question (3 pts. each).

7) The correct number of significant figures in the number 0.002320 is:

7) _____

A) 3

B) 4

C) 7

D) ambiguous

E) none of the above

- 8) Determine the answer to the following equation with correct number of significant figures: 8) _____
(4.123 × 0.12) + 24.2 = _____
A) 24.7
B) 24.695
C) 24.70
D) 25
E) none of the above
- 9) The correct prefix for the multiplier 1,000,000 is: 9) _____
A) mega.
B) micro.
C) milli.
D) nano.
E) none of the above
- 10) Which state of matter has indefinite shape and is compressible? 10) _____
A) plasma
B) liquid
C) solid
D) gas
E) none of the above
- 11) How would you classify salt water? 11) _____
A) pure substance-element
B) mixture-heterogeneous
C) mixture-homogeneous
D) pure substance-compound
E) none of the above
- 12) If a particular process is endothermic, the reverse process must be a (an) 12) _____
A) chemical change.
B) isothermal process.
C) endothermic process.
D) exothermic process.
E) none of the above
- 13) What is the value of 98 °F in units of °C? 13) _____
A) 37
B) 371
C) 22
D) 72
E) none of the above
- 14) An atom containing 7 protons, 8 neutrons, and 7 electrons 14) _____
A) is an ion.
B) cannot exist.
C) is charge-neutral.
D) is an oxygen atom.
E) none of the above

- 15) Which of the following elements has only 12 protons? 15) _____
A) O
B) C
C) Zn
D) Mg
E) none of the above
- 16) The names of the elements whose symbols are Si, P, Mn, and S are respectively, 16) _____
A) silicon, phosphorus, magnesium, and sulfur.
B) silicon, potassium, magnesium, and sulfur.
C) silicon, potassium, magnesium, and sodium.
D) silver, phosphorus, magnesium, and sulfur.
E) silicon, phosphorus, manganese, and sulfur.
- 17) Metals are located where on the periodic table? 17) _____
A) zig-zag diagonal line
B) right side
C) left side
D) middle
E) none of the above
- 18) How many protons and electrons are present in O^{2-} ? 18) _____
A) 8 protons and 10 electrons
B) 10 protons and 8 electrons
C) 8 protons and 8 electrons
D) 16 protons and 8 electrons
E) none of the above
- 19) What is the charge on the ion formed by selenium? 19) _____
A) 2-
B) 2+
C) 1-
D) 1+
E) none of the above
- 20) How many total atoms are in the formula $Al_2(CO_3)_3$? 20) _____
A) 12
B) 9
C) 8
D) 14
E) none of the above
- 21) Which among the following elements does NOT exist as a diatomic molecule in nature? 21) _____
A) neon
B) hydrogen
C) nitrogen
D) fluorine
E) none of the above

22) What is the formula for an ionic compound made of magnesium and sulfur? 22) _____
A) MgS₂
B) MgS
C) Mg₂S₃
D) Mg₂S
E) none of the above

23) What is the name of the compound made from lithium and oxygen? 23) _____
A) lithium oxide
B) oxygen lithide
C) lithium dioxide
D) lithium(I) oxide
E) none of the above

TRUE/FALSE. In scantron fill the circle "A" for a True answer and "B" for False answer (3 pts. each).

24) The mass of an object, 4.55×10^{-3} g, expressed in decimal notation is 0.000455 g. 24) _____

25) The charges on electrons and neutrons cancel each other to give neutral atoms. 25) _____

26) A cation forms when an atom gains an electron. 26) _____

27) All elements have three or more naturally occurring isotopes. 27) _____

28) The proper name for SF₆ is sulfur tetrafluoride. 28) _____